

Choose the one incorrect statement
or choose the best statement.

a. For the reaction



ionic eqn
 $\text{NH}_4^+ + \text{OH}^- + \text{H}^+ + \text{Cl}^- \rightarrow \text{NH}_4^+$
 ~~H_2O~~ Cl^-

spectator $\rightarrow \text{NH}_4^+, \text{Cl}^-$

- b. In a redox reaction: thing being reduced gains electrons (becomes smaller #), thing being oxidized loses electrons (becomes larger #).
- c. Oxidation state of an element (in its elemental state) is zero. (ex: Na, H₂ both have zero ox state)
- d. In SO₄⁻², the oxidation state of S can be calculated by the equation -2 = S + [2*(-2)]
- e. All statements are correct.

In correct
 $-2 = S + 4(-2)$