Quiz VI	General Chemistry I Lecture Fall 12 Dr. Hahn 20 pts 11/8 R (8:30) am quiz #
Name	Name
(print nam	
	ow all work for full credit.
1. a.	Give the electron configuration for the element Sn using the $1s^2$, $2s^2$ nomenclature (4 pts 25^2 , $2p^6$, 35^2 , $3p^6$, 45^7 , $3d^6$, $4p^6$, $5s^2$, $4d^{10}$, $5p^2$
b.	Give the <u>valence</u> electron configuration for the element using the same notation. (3 pts) 55^{2} , $5p^{2}$
c.	electronic of the or down errors (3 nts)
	electrons as up of down arrows. (5 pts) $\frac{1}{5}$ $\frac{1}{5}$
d.	
e.	The bigger ionization energy is for the element (circle one) (Na) or (Rb) (1 pt)
f.	The more reactive element is (circle one) (Li) or (K) (1 pt)
2.	In the p subshell, there are (give # in blank) orbitals. (2 pt)
Maxi	imum number of electrons in the p subshell is (give # in blank) (3 pts)
The	s block of the periodic table consists of Group <u>IA</u> to Group <u>IA</u> . (using the exact group er in the periodic table handed out with this quiz) (2 pts)
Extra Cre frame by:	
a.	Give the total number of valence electrons for the molecule $\frac{20}{100}$ (1 pt)
b.	Complete the Lewis Dot Structure (If you need space for a trial structure, please use the back of
	the page. You do not need to show the final structure in the space below just write in space below that the structure is shown on the back.) Hint: Lewis Dot structure has one double
	bond. Carbon cannot expand the octet. Formula for the Lewis Dot structure is: NO ₂ CH ₂ Cl (2 pts)
rialtel	(2 pts) (6 x 2 = 32
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Quiz VI General Chemistry I Lecture II) Fall 12 Dr. Hahn 20 pts 11/8 R 8:30 am quiz #
NameName(print name) Name(sign name)
Please show all work for full credit.
a. Give the electron configuration for the element x_e using the $1s^2$, $2s^2$ nomenclature (4 pts) $(5^2, 2s^2, 2p^6, 35^2, 3p^6, 4s^2, 3d^6, 4p^6, 5s^2, 4d^6, 5p^6$
b. Give the <u>valence</u> electron configuration for the element using the same notation. (3 pts) 55^2 , $5p^6$
c. Give the <u>orbital diagram</u> for the <u>valence</u> electrons of the element including showing the electrons as up or down arrows. (3 pts) 11 12 15 15
d. The bigger atomic size (atomic radius) is the element(circle one) (Li) or (Be) (1 pt)
e. The bigger ionization energy is for the element (circle one) (K) or (Ca) (1 pt)
f. The more reactive element is (circle one) (Rb) or (Li) (1 pt)
2. In the d subshell, there are (give # in blank) orbitals. (2 pt)
Maximum number of electrons in the d subshell is (3 pts)
The p block of the periodic table consists of Group IIA to Group IIIA (using the exact group number in the periodic table handed out with this quiz) (2 pts)
Extra Credit: Give the Lewis Dot Structure for the molecule for which I have provided the formula and frame by:
 a. Give the total number of valence electrons for the molecule
trial#1 101 /0x2=20\(\tilde{\text{E}}\) \\ \text{top many \(\text{E}\)} \\ \text{Styncture} \\ Stync

Name (print nam	Name (sign name)
	ow all work for full credit.
1. a.	Give the electron configuration for the element Te using the $1s^2$, $2s^2$ nomenclature (4 pts) 5^2 , 25^2 , $2p^6$, 35^2 , $3p^6$, 45^2 , $3d^{10}$, $4p^6$, 55^2 , $4d^{10}$, $5p^4$
b.	Give the <u>valence</u> electron configuration for the element using the same notation. (3 pts) 55^2 , 56^4
c.	Give the <u>orbital diagram</u> for the <u>valence</u> electrons of the element including showing the electrons as up or down arrows. (3 pts)
d.	The bigger atomic size (atomic radius) is the element(circle one) (N) or (P) (1 pt)
e.	The bigger ionization energy is for the element (circle one) (C) or (Ge) (1 pt)
f.	The more reactive element is (circle one) (K) or (Na) (1 pt)
2.	In the f subshell, there are (give # in blank) orbitals. (2 pt)
Maxi	mum number of electrons in the \mathbf{f} subshell is $\underline{////}$ (give # in blank) (3 pts)
	block of the periodic table consists of Group A to Group A . (using the exact group er in the periodic table handed out with this quiz) (2 pts)
Extra Cre frame by:	dit: Give the Lewis Dot Structure for the molecule for which I have provided the formula and
a.	Give the total number of valence electrons for the molecule $\frac{20}{100}$ (1 pt)
b.	Complete the Lewis Dot Structure (If you need space for a trial structure, please use the back of the page. You do not need to show the final structure in the space below just write in space below that the structure is shown on the back.) Hint: Lewis Dot structure has one double bond. Carbon and Nitrogen cannot expand its octet. Formula for the Lewis Dot structure is:
rial#1	101 too many kekule N-C-CII E Structure 2 0 15 x 2= 307 Kekule Structure 1 (2) 1 (2)

Name	Name (sign name)
(print nan	\mathcal{O} (sign name)
Please she	ow all work for full credit.
1. a.	Give the electron configuration for the element I using the $1s^2$, $2s^2$ nomenclature (4 pt 5^2 , 25^2 , $2p^6$, 35^2 , $3p^6$, 45^2 , $3d^{10}$, $4p^6$, 55^2 , $4d^{10}$, $5p^5$
b.	Give the <u>valence</u> electron configuration for the element using the same notation. (3 pts)
	95 ² , 5 p ⁵
c.	Give the <u>orbital diagram</u> for the <u>valence</u> electrons of the element including showing the electrons as up or down arrows. (3 pts) 1 1 1 5 p
d.	- · · · · · · · · · · · · · · · · · · ·
e.	The bigger ionization energy is for the element (circle one) (P) or (Cl) (1 pt)
f.	The more reactive element is (circle one) (Na) or (Rb) (1 pt)
_	In the p subshell, there are (give # in blank) orbitals. (2 pt)
2.	•
Maxi	mum number of electrons in the p subshell is (give # in blank) (3 pts)
	p block of the periodic table consists of Group $\frac{\text{II} \ A}{\text{II}}$ to Group $\frac{\text{VII} \ A}{\text{II}}$. (using the exanumber in the periodic table handed out with this quiz) (2 pts)
Extra Cre frame by:	
	2.7
a.	Give the total number of valence electrons for the molecule (1 pt)
b.	Give the total number of valence electrons for the molecule $2(4) + 2(7) + 2(1) = 24$ Complete the Lewis Dot Structure (If you need space for a trial structure, please use the base
	the page. You do not need to show the final structure in the space below just write in space below that the structure is shown on the back.) Hint: Lewis Dot structure has one doubt
	bond. Carbon cannot expand its octet. Formula for the Lewis Dot structure is: C_2Cl_2E
, u l	pts) 2-26 (+00 many) 142
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Quiz VI	General Chemistry I Lecture 1 Fall 12 Dr. Hahn 20 pts 11/8 R 8:30 am quiz #
Name (print nan	Name (sign name)
Please she	ow all work for full credit.
1. a.	Give the electron configuration for the element Sn using the $1s^2$, $2s^2$ nomenclature (4 pts)
b.	Give the <u>valence</u> electron configuration for the element using the same notation. (3 pts)
c.	Give the <u>orbital diagram</u> for the <u>valence</u> electrons of the element including showing the electrons as up or down arrows. (3 pts)
d.	The bigger atomic size (atomic radius) is the element(circle one) (Li) or (K) (1 pt)
e.	The bigger ionization energy is for the element (circle one) (Na) or (Rb) (1 pt)
f.	The more reactive element is (circle one) (Li) or (K) (1 pt)
2.	In the p subshell, there are (give # in blank) orbitals. (2 pt)
Maxi	mum number of electrons in the p subshell is (give # in blank) (3 pts)
	s block of the periodic table consists of Group to Group (using the exact group er in the periodic table handed out with this quiz) (2 pts)
Extra Cre frame by:	
a.	Give the total number of valence electrons for the molecule(1 pt)
b.	Complete the Lewis Dot Structure (If you need space for a trial structure, please use the back of the page. You do not need to show the final structure in the space below just write in space below that the structure is shown on the back.) Hint: Lewis Dot structure has one double bond. Carbon cannot expand the octet. Formula for the Lewis Dot structure is: NO ₂ CH ₂ Cl (2 pts)
	О
	N C Cl
0	T

Name	Name
(print nam	e) (sign name)
Please sho	w all work for full credit.
1. a.	Give the electron configuration for the element \mathbf{Xe} using the $1s^2, 2s^2 \dots$ nomenclature (4 pts)
b.	Give the <u>valence</u> electron configuration for the element using the same notation. (3 pts)
c.	Give the <u>orbital diagram</u> for the <u>valence</u> electrons of the element including showing the electrons as up or down arrows. (3 pts)
d.	The bigger atomic size (atomic radius) is the element(circle one) (Li) or (Be) (1 pt)
e.	The bigger ionization energy is for the element (circle one) (K) or (Ca) (1 pt)
f.	The more reactive element is (circle one) (Rb) or (Li) (1 pt)
2.	In the d subshell, there are (give # in blank) orbitals. (2 pt)
Maxir	num number of electrons in the d subshell is (give # in blank) (3 pts)
	block of the periodic table consists of Group to Group (using the exact group r in the periodic table handed out with this quiz) (2 pts)
Extra Cred frame by:	lit: Give the Lewis Dot Structure for the molecule for which I have provided the formula and
a.	Give the total number of valence electrons for the molecule(1 pt)
ъ.	Complete the Lewis Dot Structure (If you need space for a trial structure, please use the back of the page. You do not need to show the final structure in the space below just write in space below that the structure is shown on the back.) Hint: Lewis Dot structure has one double bond. Carbon cannot expand octet. Formula for the Lewis Dot structure is: Cl C O H (2 pts)
	O
Cl	С Н

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Name (print nan					Name (sign_na	me)				
	ŕ				(6181.)				
Please she	ow al	l work f	for full credit.							
1. a.	Giv	e the el	ectron configuratio	n for the	element '	Te using	the $1s^2$, $2s^2$	² nome	enclature (4 pts)	
b.	Gi	ive the <u>v</u>	valence electron co	nfiguratio	on for the	element u	sing the sa	me notati	on. (3 pts)	
c.			rbital diagram for s up or down arrow			ns of the	element in	cluding sh	nowing the	
d.	The	e bigger	atomic size (atomi	c radius)	is the elen	nent(circle	e one) (N) or (P) (l pt)	
e.	The	e bigger	ionization energy	is for the	element (c	circle one)	(C) or (C	Ge) (1 pt)		
f.	The	more r	reactive element is	(circle on	e) (K) or	(Na) (1 j	ot)			
2.	In t	he f	subshell, there are	:	_(give#iı	n blank) (orbitals. (2	pt)		
Maxi	mum	numbe	r of electrons in the	e f sub	shell is		(give # i	n blank)	(3 pts)	
The s	s blo er in	ock of the	ne periodic table co odic table handed o	nsists of Cout with the	Group his quiz) (to G1 (2 pts)	oup	(using the	e exact group	
Extra Cre frame by:		Give th	ne Lewis Dot Struc	ture for th	ne molecul	le for whi	ch I have p	rovided t	he formula and	
a.	Giv	e the to	tal number of valer	nce electr	ons for the	e molecule	e	(1	pt)	
b.	the belo bor	page. ow that	he Lewis Dot Struct You do not need to the structure is sho rbon and Nitroger O (2 pts)	show the	e final stru e back.) H	cture in th lint: Le v	ne space be wis Dot str	low just v ucture h	write in space as one double	
		О								
Н	N	C	Cl							
	Cl									

Name	Name
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Please sho	ow all work for full credit.
1. a.	Give the electron configuration for the element I using the $1s^2$, $2s^2$ nomenclature (4 pts)
b.	Give the <u>valence</u> electron configuration for the element using the same notation. (3 pts)
c.	Give the <u>orbital diagram</u> for the <u>valence</u> electrons of the element including showing the electrons as up or down arrows. (3 pts)
d.	The bigger atomic size (atomic radius) is the element(circle one) (N) or (F) (1 pt)
e.	The bigger ionization energy is for the element (circle one) (P) or (Cl) (1 pt)
f.	The more reactive element is (circle one) (Na) or (Rb) (1 pt)
2.	In the p subshell, there are (give # in blank) orbitals. (2 pt)
Maxi	mum number of electrons in the p subshell is (give # in blank) (3 pts)
The group	p block of the periodic table consists of Group to Group (using the exact number in the periodic table handed out with this quiz) (2 pts)
Extra Creeframe by:	dit: Give the Lewis Dot Structure for the molecule for which I have provided the formula and
a.	Give the total number of valence electrons for the molecule (1 pt)
b.	Complete the Lewis Dot Structure (If you need space for a trial structure, please use the back of the page. You do not need to show the final structure in the space below just write in space below that the structure is shown on the back.) Hint: Lewis Dot structure has one double bond. Carbon cannot expand its octet. Formula for the Lewis Dot structure is: C ₂ Cl ₂ H ₂ (2 pts)
Cl	С С
Н	Cl