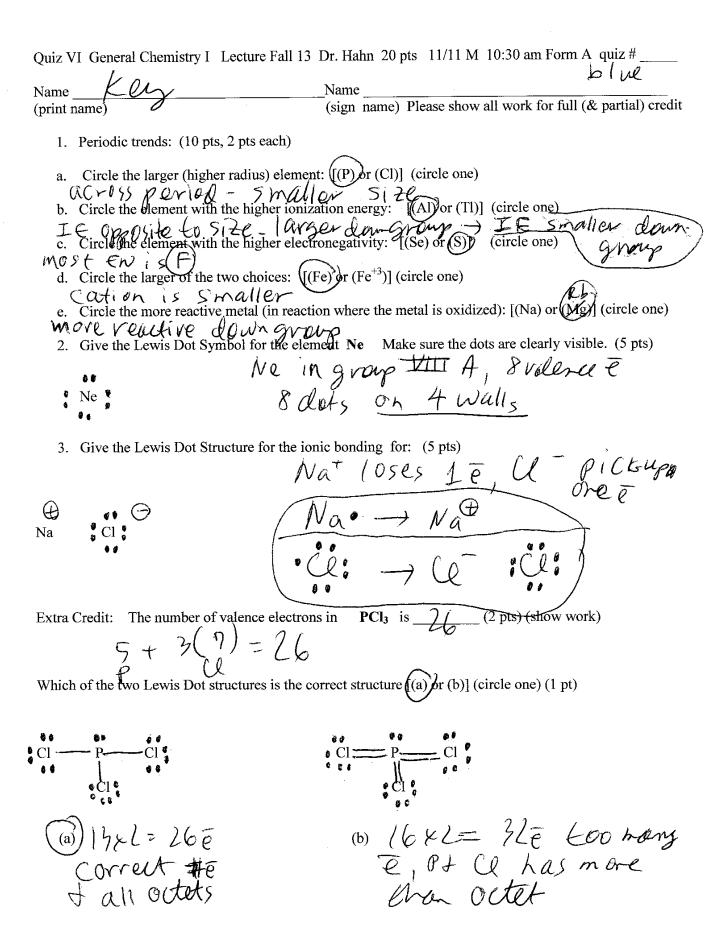
Quiz VI General Ch	nemistry I Lecture Fall	13 Dr. Hahn 20 pts 1	1/11 M 9:30	am Form A quiz#
Name(print name)	y	Name	show all worl	for full (& partial) credit
(print name)	V	(sign name) i lease	SHOW all Woll	c for full (ac partial) croar
	ds: (10 pts, 2 pts each)		1	
a. Circle the la	rger (higher radius) eleme	ent: [(Si) or (Sn)] (cir	rcle one) (in Erand pid
b. Circle the ele	ement with the higher ion	ization energy: [(K) or	(Ca) (circle	one)
c. Circle the ele	enent with the higher elec	etronegativity: [(C) o	r (N) (circle	one)
~ \	ger of the two choices:			
e. Circle the m	Smaller ore reactive metal (in reac	ction where the metal is	oxidized): [(0	Ca) or (Be)] (circle one)
(d) run	Grapp - In or vis Dot Symbol for the ele	o Veaux VE)	•	
2. Of ve the Lev	-			
• F •		Valence e	, 7 dets	, 4 walls
/ /-	2p ⁶ , 35 ² , 3p ⁶ , 45		4 -se/v	regative horye-
Se-L	-> [Ar] 45%	bd", 4p6)		lore e
			₽.	
	number of valence electro		$\underline{\mathcal{S}}$ (2 pts)	(show work)
valence	$\bar{e} = 2(1)$	+ 6 = 8e		
Which of the two L	ewis Dot structures is the	correct structure ((a))or	(b)] (circle on	e) (1 pt)
			. / .	, , ,
, * ⊕		4 *		
	//			
н н	H	H		
(a) 4 x 2	$= \mathcal{F}\bar{e}$ (b)	9x1 = 18x	5 COAM	Sam O
	L+	9x2=180	h 10 1). A
	1(Carrie 1	gue de	uel

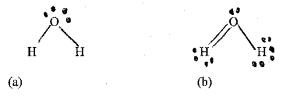
Name /<	Ch/	Name
(print name)	8	(sign name) Please show all work for full (& partial) cred
1 Period	lic trends: (10 pts, 2 pts e	each)
	, <u>-</u> -	
a. Circle	e the larger (higher radius	as) element: ((Sr) or (Be)] (circle one)
b. Circle	the element with the high	gher ionization energy: [(C) or (O)] (circle one)
IE	- opposite to	512e - across porgo smaller Size bu
c. Circle	the element with the high	gher ionization energy: [(C) or (O)) (circle one) S 20 - across for all smaller 5/20 but gher electronegativity: [(Br) or (Cl)] (circle one) a - yer I
d. Circle	the larger of the two cho	oices: [(Cl) or (Cl ⁻¹)] (circle one)
N O.	wating in	(in reaction where the metal is oxidized): [(Li) or (K)] (circle one)
e. Circle	the more reactive metal	(in reaction where the metal is oxidized). [(El) of (R)] (energ one)
2. Give t	•	or the element Si Make sure the dots are clearly visible. (5 pts)
		Si in group # A - 4 Vollacle
• Si •		- lullat
V		12 por each of
3. Give t	the electron configuration	on for the ion shown: Ca+2 (5 pts) 1 Per each of 4 (5 pts) 4 walls
1	c2 252 2p6	357,386, 452) - Ca 7 Walls
	OV ([Ar]	1/4.62
(+)	260-20 19 11 2 2	- / 1
1	- vamoul 2e	LAN or L (2t
	$(15^2)_{5^2}$	206 35 30E) T
Evtra Credit:	The number of valence	re electrons in CCl ₄ is
C	(l	(2 pts) (site 11 visit)
4+	4(7) = 31	_
Which of the		es is the correct structure [(a) or (b)] circle one) (1 pt)
	4.9	
	ici Cincorre	ELF CI
• C1		
0 00	00	\$ 00
	ci ;	ू <u>C</u> 1 है
· a.	21 - fr	Domony (b) 16x2=32e none C notable be N=2 has no d qub
$/$ (a) $ \mathcal{Y} $	xL - 166 9	(b) 16×L= 76
	J Cl has m	Chalant MET KAS NO & SIND



(sign name) Please show all work for full (& partial) of the Periodic trends: (10 pts, 2 pts each) a. Circle the larger (higher radius) element: ((Ga) or (Br)] (circle one) \$\sin \text{12} (le CV like \(\text{	Name Cly	Name
a. Circle the larger (higher radius) element: [(Ga) or (Br)] (circle one) \$\frac{1}{2} \therefore (\therefore \therefore \there \therefore \there \therefore \there \therefore	(print name)	Name (sign name) Please show all work for full (& partial) contains the same of the same o
a. Circle the larger (higher radius) element: ((Ga) or (Br)] (circle one) SIZE (PEN LOLE) ACNOSS PEN BA Circle the element with the higher ionization energy: (Ba) por (Be) (circle one) LE OPPOLITE TO SIZE (DUMAN GOS) Circle the element with the higher electronegativity: [(SiJ or (S))] (circle one) Circle the larger of the two choices: [(S) or (S)) (circle one) ANION IS LIGHT C. Circle the more reactive metal (in reaction where the metal is oxidized): [(K) or (Da)] (circle one) 2. Give the Lewis Dot Symbol for the element In Make sure the dots are clearly visible. (5 pts) In ON 4 WALLS 3. Give the Lewis Dot Structure for the covalent bonding for: (5 pts) UN Lorded BY XX Extra Credit: The number of valence electrons in NH3 is (2 pts) (show work) 5 + 3 (i) = 8 Which of the two Lewis Dot structures is the correct structure [(a) o(b)) (circle one) (1 pt) (NOWAN H N H	1. Periodic trends: (10 pts, 1	2 pts each)
b. Circle the element with the higher ionization energy: 'Balant (Be) (circle one) 1	-	
Circle the element with the higher electronegativity: [(Si) o ((S)) (circle one) P S MOST WAS MOST WA	a. Circle the larger (higher)	radius) element: [(Gayor (Br)] (circle one)
e. Circle the larger of the two choices: [(S) or (S ²) (circle one) e. Circle the more reactive metal (in reaction where the metal is oxidized): [(K) or (Ea)] (circle one) e. Circle the more reactive metal (in reaction where the metal is oxidized): [(K) or (Ea)] (circle one) ? A way where 2. Give the Lewis Dot Symbol for the element In Make streethe dots are clearly visible. (5 pts) In On 4 walls 3. Give the Lewis Dot Structure for the covalent bonding for: (5 pts) Where a better the dots are clearly visible. (5 pts) In On 4 walls Extra Credit: The number of valence electrons in NH ₃ is (2 pts) (show work) The structure is the correct structure [(a) o(b)) (circle one) (1 pt) (MC ON A H H H H H H H H H H H H H H H H H H	b. Circle the element with the	he higher ionization energy: (Ba) por (Be) (circle one)
e. Circle the larger of the two choices: [(S) or (S ²) **Circle one) e. Circle the more reactive metal (in reaction where the metal is oxidized): [(K) or (Ea)] (circle one) 2. Give the Lewis Dot Symbol for the element In Make streethe dots are clearly visible. (5 pts) In On 4 Wills 3. Give the Lewis Dot Structure for the covalent bonding for: (5 pts) When he had Br XX Extra Credit: The number of valence electrons in NH ₃ is (2 pts) (show work) The had been declarated by the correct structure [(a) or (b)) **Circle one) (1 pt) (NOWER H. H	IE Opposite to	5/2e downgroup Rigger S, Ze, >m
e. Circle the more reactive metal (in reaction where the metal is oxidized): [[K] or [Pa]] (circle one reactive in the more reactive metal (in reaction where the metal is oxidized): [[K] or [Pa]] (circle one reactive in the more reactive metal (in reaction where the metal is oxidized): [[K] or [Pa]] (circle one reactive in the content of the content in the metal is oxidized): [[K] or [Pa]] (circle one reactive in the content in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal is oxidized): [[K] or [Pa]] (circle one in the reactive in the metal in the m	Fis mast EN	the higher electronegativity. [(31) of (3)] (ellele one)
2. Give the Lewis Dot Symbol for the element In Make sure the dots are clearly visible. (5 pts In A Walls 3. Give the Lewis Dot Structure for the covalent bonding for: (5 pts) Charles Br Br Br Br Br Credit: The number of valence electrons in NH3 is P (2 pts) (show work) 5 +3(1) = 8 Which of the two Lewis Dot structures is the correct structure [(a) o(b)) (circle one) (1 pt) IN CAPUAL H COVEAL H IN H I	deircle the larger of the tw	wo choices: [(S) or (S ⁻²)] (circle one)
2. Give the Lewis Dot Symbol for the element In Make sure the dots are clearly visible. (5 pts In A Walls 3. Give the Lewis Dot Structure for the covalent bonding for: (5 pts) Un bonded Br XX Br Br Ax Extra Credit: The number of valence electrons in NH3 is P (2 pts) (show work) 5 +3(1) = 8 Which of the two Lewis Dot structures is the correct structure [(a) o((b))) (circle one) (1 pt) IN COVERT H IN	e Circle the more reactive t	metal (in reaction where the metal is oxidized): [(K) or (2a)] (circle one
In on 4 wills 3. Give the Lewis Dot Structure for the covalent bonding for: (5 pts) Un bonded Br Extra Credit: The number of valence electrons in NH3 is	reactivity inc	reases down group
3. Give the Lewis Dot Structure for the covalent bonding for: (5 pts) Un bonded Br Br X Br X XX Extra Credit: The number of valence electrons in NH3 is (2 pts) (show work) 5 +3(1) = 8 Which of the two Lewis Dot structures is the correct structure [(a) or (b)] (circle one) (1 pt) (N C O V E A H N H N H H H N H H H N H N H H N H H N H H N H H N H H N H H N H H N H		
3. Give the Lewis Dot Structure for the covalent bonding for: (5 pts) Un borded Br Br X Br X XX Extra Credit: The number of valence electrons in NH3 is (2 pts) (show work) 5 +3(1) = 8 Which of the two Lewis Dot structures is the correct structure [(a) of (b)] (circle one) (1 pt) IN COVERT H N N		-n) in group IIIH, 3 valence e, 3 dot,
Extra Credit: The number of valence electrons in NH ₃ is	• In	on 4 walls
Extra Credit: The number of valence electrons in NH ₃ is	•	
Extra Credit: The number of valence electrons in $\mathbb{N}H_3$ is \mathbb{R} (2 pts) (show work) $ \begin{array}{c} 5 + 3(1) = 8 \\ \mathbb{N} \\ \mathbb$	3. Give the Lewis Dot Struc	cture for the covalent bonding for: (5 pts)
Extra Credit: The number of valence electrons in $\mathbb{N}H_3$ is \mathbb{R} (2 pts) (show work) $ \begin{array}{c} 5 + 3(1) = 8 \\ \mathbb{N} \\ \mathbb$	Unbord	led Br xx
Extra Credit: The number of valence electrons in NH_3 is		- So + x Bx
Extra Credit: The number of valence electrons in NH_3 is	Br & Br X	* XX
Which of the two Lewis Dot structures is the correct structure $[(a) \circ (b)]$ (circle one) (1 pt) (NCOVELL H N H H H H	0 # XX	
Which of the two Lewis Dot structures is the correct structure $[(a) \circ (b)]$ (circle one) (1 pt) (NCOVELL H N H H H H		\sim
Which of the two Lewis Dot structures is the correct structure [(a) of (b)] (circle one) (1 pt) (N C OVEN H N H H H H	_	
Which of the two Lewis Dot structures is the correct structure [(a) of (b)] (circle one) (1 pt) (N COVENT H N H H H H	5 +3(1) =	= 8
incorrect Correct H—N—H H H H	NH	
	Which of the two Lewis Dot stru	uctures is the correct structure [(a) or (b)] (circle one) (1 pt)
	the change to	C 2 2 2 2 4
H N H	•	Correct
	•H = N _ H	$H \longrightarrow N \longrightarrow H$
		<u>.l</u>
(a) 13 x 2=26 = t (b) 4 x 2= 8 =, H has toomang et duet, N has outet	TH C	н .
(a) 15 X L= 26c + (b) 4 x 2 - 00, H Mas outet	12.1.2 2/=	(1 / 1 - A = 11 has
(Evomony e duet, Nhas ottet	(a) 15 X L= 26 e	(b) 4 × 2 - 0 e H NOS
	1 toomange	duet, Nhas other
I He cannot have	1 11 carnet NO	We Lawrest # 5

Quiz VI General Chemistry I Lec	Name
(print name)	(sign name) Please show all work for full (& partial) cred
1. Periodic trends: (10 pts, 2 p	ots each)
a. Circle the larger (higher rad	dius) element: [(Si) or (Sn)] (circle one)
b. Circle the element with the	higher ionization energy: [(K) or (Ca)] (circle one)
c. Circle the element with the	higher electronegativity: [(C) or (N)] (circle one)
d. Circle the larger of the two	choices: [(Na) or (Na ⁺)] (circle one)
e. Circle the more reactive me	tal (in reaction where the metal is oxidized): [(Ca) or (Be)] (circle one)
2. Give the Lewis Dot Symbo	for the element F Make sure the dots are clearly visible. (5 pts)
F	
3. Give the electron configura	tion for the ion shown: Se ⁻² (5 pts)

Which of the two Lewis Dot structures is the correct structure [(a) or (b)] (circle one) (1 pt)



Name	Name		
	(sign	name)	Please show all work for full (& partial) credit

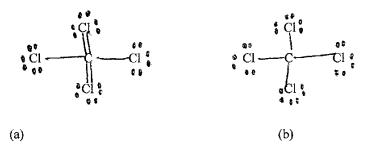
- 1. Periodic trends: (10 pts, 2 pts each)
- a. Circle the larger (higher radius) element: [(Sr) or (Be)] (circle one)
- b. Circle the element with the higher ionization energy: [(C) or (O)] (circle one)
- c. Circle the element with the higher electronegativity: [(Br) or (Cl)] (circle one)
- d. Circle the larger of the two choices: [(Cl) or (Cl⁻¹)] (circle one)
- e. Circle the more reactive metal (in reaction where the metal is oxidized): [(Li) or (K)] (circle one)
- 2. Give the Lewis Dot Symbol for the element Si Make sure the dots are clearly visible. (5 pts)

Si

3. Give the electron configuration for the ion shown: Ca^{+2} (5 pts)

Extra Credit: The number of valence electrons in CCl₄ is _____ (2 pts) (show work)

Which of the two Lewis Dot structures is the correct structure [(a) or (b)] (circle one) (1 pt)



Name _ (print r	Name
1.	Periodic trends: (10 pts, 2 pts each)
a.	Circle the larger (higher radius) element: [(P) or (Cl)] (circle one)
b.	Circle the element with the higher ionization energy: [(Al) or (Tl)] (circle one)
c.	Circle the element with the higher electronegativity: [(Se) or (S)] (circle one)
d.	Circle the larger of the two choices: [(Fe) or (Fe ⁺³)] (circle one)
e.	Circle the more reactive metal (in reaction where the metal is oxidized): [(Na) or (Mg)] (circle one)
2.	Give the Lewis Dot Symbol for the element Ne Make sure the dots are clearly visible. (5 pts)
	Ne
3.	Give the Lewis Dot Structure for the ionic bonding for: (5 pts)
Na	Cl
Extra	Credit: The number of valence electrons in PCI ₃ is (2 pts) (show work)
Which	n of the two Lewis Dot structures is the correct structure [(a) or (b)] (circle one) (1 pt)
Cl	P—————————————————————————————————————

(b)

(a)

Name	NameName
(print name)	(sign name) Please snow all work for full (& partial) credit
1. Periodic trends: (10 p	pts, 2 pts each)
a. Circle the larger (hig	gher radius) element: [(Ga) or (Br)] (circle one)
b. Circle the element wi	ith the higher ionization energy: [(Ba) or (Be)] (circle one)
c. Circle the element wi	ith the higher electronegativity: [(Si) or (S)] (circle one)
d. Circle the larger of th	he two choices: [(S) or (S ⁻²)] (circle one)
e. Circle the more react	tive metal (in reaction where the metal is oxidized): [(K) or (Ca)] (circle one)
2. Give the Lewis Dot S	Symbol for the element In Make sure the dots are clearly visible. (5 pts)
In	
3. Give the Lewis Dot S	Structure for the covalent bonding for: (5 pts)
Br Br	
Extra Credit: The number	of valence electrons in NH ₃ is(2 pts) (show work)
Which of the two Lewis Dot	ot structures is the correct structure [(a) or (b)] (circle one) (1 pt)
•H —N — H •	H—————————————————————————————————————