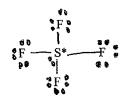
Name		_
(print r	name) (sign name) e show work for full credit and to get partial credit.	
		-
1.	For the following Lewis Dot structure, complete the following.	
,		
g B	-S-12	
8F		
4	F	
	How many electron pairs (VSEPRT electron pair) are on the atom with the *?	2 nt
	1	r·
b.	How many lone pairs are on the atom with the *?3 pt	
	What is the the geometry of the electron pairs? Evigaral biph vami	dúl3 pt
d.	What is the geometry of the molecule? See Saw	3 pt
	What is the hybridization of the atom with the *? $\frac{5\rho^3d}{}$	3 pt
٠.	1200 000	
f.	What is the bond angle in the molecule? $\frac{120^{\circ}}{90^{\circ}}$	3 pt
	· ·	to the Lewis Det
2.	Draw the molecule above in 3D using the wedge, dash, line drawing in the space next structure. Remember that (coming away from plane) (going below (in plane) (3 pt)	w plane (1999)
2. Extra (Draw the molecule above in 3D using the wedge, dash, line drawing in the space next structure. Remember that (coming away from plane) (going below	w plane (1999)
2. Extra (Draw the molecule above in 3D using the wedge, dash, line drawing in the space next structure. Remember that (coming away from plane) (going below (in plane) (3 pt) Credit (3 pts, 1 pt per letter) a. For the bond between the atoms show the dipole	w plane (1999)
Extra 6 the fol	Draw the molecule above in 3D using the wedge, dash, line drawing in the space next structure. Remember that (coming away from plane (in plane) (3 pt) (3 pt) Credit (3 pts, 1 pt per letter) a. For the bond between the atoms show the dipole $C = C = C = C = C = C = C = C = C = C $	w plane (1999)
2. Extra (Draw the molecule above in 3D using the wedge, dash, line drawing in the space next structure. Remember that (coming away from plane) (going below (in plane) (3 pt) Credit (3 pts, 1 pt per letter) a. For the bond between the atoms show the dipole lowing +> or the δ +/ δ - CCl For the 3D drawing below, show all individual dipole moment arrows.	w plane (1999)
Extra 6 the fol	Draw the molecule above in 3D using the wedge, dash, line drawing in the space next structure. Remember that (coming away from plane) (going below (in plane) (3 pt) Credit (3 pts, 1 pt per letter) a. For the bond between the atoms show the dipole lowing +> or the δ +/ δ - CCl For the 3D drawing below, show all individual dipole moment arrows.	w plane (1999)
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Extra the following the follow	Draw the molecule above in 3D using the wedge, dash, line drawing in the space next structure. Remember that (coming away from plane (in plane) (3 pt) Credit (3 pts, 1 pt per letter) a. For the bond between the atoms show the dipole lowing +> or the δ +/ δ - C	w plane (1994)
Extra 6 the fol	Draw the molecule above in 3D using the wedge, dash, line drawing in the space next structure. Remember that (coming away from plane (in plane) (3 pt) (going below (in plane) (3 pt) Credit (3 pts, 1 pt per letter) lowing +> or the δ+/δ- For the 3D drawing below, show all individual dipole moment arrows. Cl + 0 - different Is the molecule above [(polar) or (nonpolar)] (circle one of the two parenthesis) Explant	w plane (1994)
Extra the following the follow	Draw the molecule above in 3D using the wedge, dash, line drawing in the space next structure. Remember that (coming away from plane (in plane) (3 pt) Credit (3 pts, 1 pt per letter) a. For the bond between the atoms show the dipole lowing +> or the δ +/ δ - C	w plane (1994)
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Extra the following the follow	Draw the molecule above in 3D using the wedge, dash, line drawing in the space next structure. Remember that (coming away from plane (in plane) (3 pt) (going below (in plane) (3 pt) Credit (3 pts, 1 pt per letter) lowing +> or the δ+/δ- For the 3D drawing below, show all individual dipole moment arrows. Cl + 0 - different Is the molecule above [(polar) or (nonpolar)] (circle one of the two parenthesis) Explant	w plane (1994)

Name KM	Name	(green
(print name)	(sign name)	
Please show work for full cred	it and to get partial credit.	00
1 Partha Collarying Lavin	Dot structure, complete the following.	(all same)
1. For the following Lewis	Dot structure, complete the following.	n/and)
	101	piace
S O S	Y	
	چ ع ر	- 1)-4
C H		-N
, 3 ¢	1, 0	X b
A		3
a. How many electron pair	rs (VSEPRT electron pair) are on the atom with	the * ?2 pt
b. How many lone pairs a	re on the atom with the *?3 pt	
		1 1-
 c. What is the the geometre 	ry of the electron pairs?tvigona	1 planav 3 pt
d What is the asometry o	fthe molecule? + 1/16 And 0/10	3 pt 3 pt 3 pt
u. What is the geometry of	f the molecule? tvigoral pla	- F
e. What is the hybridization	on of the atom with the *?	3 pt
	1 ~ ^0	3 pt
f. What is the bond angle	in the molecule !	
2. Draw the molecule abo	ve in 3D using the wedge, dash, line drawing in	the space next to the Lewis Dot
structure. Remember 1) (going below plane IIIIII)
(in plane	-) $(3pt)$	
` - /	For the bond between the atoms show the dipole	
+> or the δ + / δ -	 	(pt. por)
	XeF	(pet per)
b. For 3D drawing below,	show all individual dipole moment arrows.	1 evies
200	·	
F /// OC XF		
F X X YF	\	
Ta 4h a 22 - 1 1 1	[(-1	anthogis) Evalgin
	[(polar) or (nonpolar)] (circle one of the two particle) V_1 V_2 V_3 V_4 V_5 V_6 V	,
	~ /	4 011 04

Name Ken	Name	· ·
(print name)	(sign name)	
Please show work for full credit as		
	1 . 4 . 6.11	
1. For the following Lewis Do	at structure, complete the following.	
% 4 ⊅ C I &		
Cl — \$b*——Cl		
6 6 6 6 6 6 6 6 6 6 6 6 6 6 6 6 6 6 6		
-		(204)
a. How many electron pairs (VSEPRT electron pair) are on the atom with t	he *?
b. How many lone pairs are o	n the atom with the *? (3pt)	
		$\alpha = \alpha = \alpha$
c. What is the the geometry of	f the electron pairs? GCHUCK	(5pt)
J What is the groundtwy of th	a malagula 2 SQ 1/0mg Q	vanidal()
d. what is the geometry of the	e molecule? Square po	(3/2)
	of the atom with the *?	(30t)
	000	(3)
f. What is the bond angle in	the molecule?	(Gpt)
2. Draw the molecule above i	in 3D using the wedge, dash, line drawing in t	he space next to the Lewis Dot
structure. Remember that	(coming away from plane) (going below plane ((())
(in plane	-) $(3pt)$	
Extra Credit (3 pts) a. For	the bond between the atoms show the dipole	of the bond using the following
$+>$ or the $\delta+/\delta-$	- -	Clatra
+		() pr per
C	Cl	(sher)
4	t ml-	· 01.00
b. For 3D drawing below, she	ow all individual dipole moment arrows.	
н	Ct Halanten	
Ĩ,	o covered yame	
XS TITLES	electro negativo	TZ
Classic	0	0

Name	10m	Name	gre
(print name)	work for full credit a	(sign name) nd to get partial credit.	
1. For the	e following Lewis D	ot structure, complete the followin	g.
, P.*		- all insame	plane
0.3	Ò		
			atom with the *? $\frac{3}{2}$
b. How	nany lone pairs are o	on the atom with the *?	can stare
c. What	is the the geometry of	e molecule?	goral planer (3
e. What	is the bond angle in t	of the atom with the $*?$ the molecule? 120°	
			lrawing in the space next to the Lev
	are. Remember that	-	
Extra Credit ((3 pts) a. For $\theta = \delta + /\delta$	the bond between the atoms show	the dipole of the bond using the following t
		 PCl	() pt pe
	I	7 6-	CICHER
b. For 31	O drawing below, she	ow all individual dipole moment a	
F		F more EN.	
Cl	Cle	but F-P	- 1= Symmetric) and Stymme Vector Sum =
T.		Jan Ce) are stymne
4 1	- ar	1000-0	1 0



- a. How many electron pairs (VSEPRT electron pair) are on the atom with the *? _____2 pt
- b. How many lone pairs are on the atom with the *? _____3 pt
- c. What is the the geometry of the electron pairs?
- d. What is the geometry of the molecule?
- e. What is the hybridization of the atom with the *?
- f. What is the bond angle in the molecule?
- 2. Draw the molecule above in 3D using the wedge, dash, line drawing in the space next to the Lewis Dot structure. Remember that (coming away from plane) (going below plane (it the plane) (3 pt)

Extra Credit (3 pts, 1 pt per letter) a. For the bond between the atoms show the dipole of the bond using the following +---> or the δ + / δ --

C-----Cl

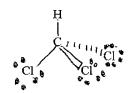
b. For the 3D drawing below, show all individual dipole moment arrows.

c. Is the molecule above [(polar) or (nonpolar)] (circle one of the two parenthesis) Explain.

Name _	Name	_
print n		
1.	For the following Lewis Dot structure, complete the following.	
e C	C H	
a.	How many electron pairs (VSEPRT electron pair) are on the atom with the *?	_2 pt
b.	How many lone pairs are on the atom with the *?3 pt	
c.	What is the the geometry of the electron pairs?	3 pt
d.	What is the geometry of the molecule ?	3 pt
e.	What is the hybridization of the atom with the *?	
f.	What is the bond angle in the molecule?	3 pt
2.	Draw the molecule above in 3D using the wedge, dash, line drawing in the space next to structure. Remember that (coming away from plane) (going below (in plane)	o the Lewis Dot
	Credit (3 pts) a. For the bond between the atoms show the dipole of the bond usi or the $\delta+/\delta-$ XeF	ng the following
ь.	For 3D drawing below, show all individual dipole moment arrows.	
F // /	Xe F	

c. Is the molecule above [(polar) or (nonpolar)] (circle one of the two parenthesis) Explain.

b. For 3D drawing below, show all individual dipole moment arrows.



c. Is the molecule above [(polar) or (nonpolar)] (circle one of the two parenthesis) Explain.

Is the molecule above [(polar) or (nonpolar)] (circle one of the two parenthesis) Explain.