	I General Chemistry I Lecture Spring 15 Dr. Hahn 20 pts 4/21 T 8:30am form A quiz #
Vame _ print na	Name (sign name)
Please s 10:30an	how all work for full credit and for partial credit. Final Exam 4/30 8:30 am Tues, Thurs class 8:30- n; 9:55 am Tues, Thurs class 3-4pm in LSF 301 1/3 multiple choice, 1/3 short answer, 1/3 long answer
1.	a. Give the Lewis Dot symbol of the element N (4 pts)
	b. Number of valence electrons for the element above is: electrons (4 pts)
2.	Give the Lewis Dot Structure of the following by completing the following. HNOCI +1
2.0	Show your work for the count of the valence electrons in the entire molecules. (6 pts)  H  O  (1e) + (5e) + (6e) + 7e - 1e = (8e)
	는 이 수 있는 사람들은 그리고 있는 것 같아요. 그리고 있는 것이 되었다면 보고 있는 것 같아 보고 있다면 보다 보고 있다면 보고 되었다면 보고 있다면
b.	Given the following two structures, choose the correct structure and then explain one reason why the <u>other</u> structure is <u>incorrect.</u> (6 pts)
н	(a) convention (b) Per (2) H can anhydrate duct
7 Extra C	Pair X2=18E  (2) N in 2nd period Cant  Credit (3 pts, ½ pt each blank) For the following given Lewis Dot Structure, complete the following:  Expand other  H  H
a.	Number of electrons pairs on the atom with the * for VSEPRT
h	Number of lone pairs on the atom with the * 1
c.	Geometry of the electron pairs at the atom with the * te Grahedral
d.	Geometry of the electron pairs at the atom with the * trigoral pyramidal
e.	Bond angle at the atom with the * 109.5°
	Hybridization at the atom with the * $\frac{5p^3}{}$

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Please show all work for full cred 10:30am; 9:55 am Tues, Thurs cla	lit and for partial credit. <u>Final Exam 4/30</u> 8:30 am Tues, Thurs class 8:30-ass 3-4pm in <b>LSF 301</b> 1/3 multiple choice, 1/3 short answer, 1/3 long answer
1. a. Give the Lewis Do	ot symbol of the element Cl (4 pts)
b. Number of valence	e electrons for the element above is: electrons (4 pts)
2. Give the Lewis Dot Struct	ture of the following by completing the following. H <sub>3</sub> N C Cl <sub>2</sub>
	ount of the valence electrons in the entire molecules. (6 pts)
3 (1e) + (5	(e) + (4ē) + 2(?ē) = 26ē
H N	C
b. Given the following two the other structure is inco	structures, choose the correct structure and then explain one reason why
	orrect. (6 pts) only do dust H
H N C CI & Cha	c has more  an octet-Cin  period  Period  H  Con  Con  Con  Con  Con  Con  Con
) (6 E pair X2 =	32 E £00 mors e 13 x 2 = 2 & e  nk) For the following given Lewis Dot Structure, complete the following:
***	,
H C*	
	2
	on the atom with the * for VSEPRT
b. Number of lone pairs on the	
c. Geometry of the electron p	pairs at the atom with the * trigoral Planar
d. Geometry of the molecule	pairs at the atom with the * $trigoral Planar$ e at the atom with the * $trigoral Planar$ ith the * $120^{\circ}$
e. Bond angle at the atom with	ith the * $\sqrt{20^{\circ}}$
f. Hybridization at the atom	with the * $\frac{2}{\sqrt{5p^2}}$

20 ptg 4/21 T 9.55 am form A guiz#
Quiz VII General Chemistry I Lecture Spring 15 Dr. Hahn 20 pts 4/21 T 9:55 am form A quiz#
Name Name (sign name)
Please show all work for full credit and for partial credit. <u>Final Exam 4/30</u> 8:30 am Tues, Thurs class 8:30-10:30am; 9:55 am Tues, Thurs class 3-4pm in <b>LSF 301</b> 1/3 multiple choice, 1/3 short answer, 1/3 long answer
1. a. Give the Lewis Dot symbol of the element & C e (4 pts)
b. Is the bond NO [(polar covalent) or (non polar covalent)] (circle one) (4 pts)
2. Give the Lewis Dot Structure of the following by completing the following. H <sub>2</sub> C S O
a. Show your work for the count of the valence electrons in the entire molecules. (6 pts)
1 C S O 1 S 0 2 (1ē) + (4ē) + 6ē + 6ē = 18ē
b. Given the following two structures, choose the correct structure and then explain one reason why the other structure is incorrect. (6 pts)
$\overline{a}$
is of the one of the correct
H Cannot expand H O H  (a)  Octet  (b)
H Cannot expand "
(a) OCTet (b)
210x2=20E
Extra Credit (3 pts, ½ pt each blank) For the following given Lewis Dot Structure, complete the following:
°° CIH
C=C*
El- 1
a. Number of electrons pairs on the atom with the * for VSEPRT
b. Number of lone pairs on the atom with the * Zev 0
c. Geometry of the electron pairs at the atom with the * tvigoral pland
c. Geometry of the electron pairs at the atom with the * <u>tvigoral plands</u> d. Geometry of the molecule at the atom with the * <u>tvigoral plands</u>
e. Bond angle at the atom with the * 1 C O
f. Hybridization at the atom with the *

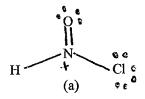
Name	Key	Name	#1 - 1
(print name)		(sign name)	
Please show a 10:30am; 9:5	all work for full credit at 5 am Tues, Thurs class 3	and for partial credit. Final Exam 4/30 8:30 -4pm in LSF 301 1/3 multiple choice, 1/3 sh	am Tues, Thurs class 8:30- ort answer, 1/3 long answer
1. a.	Give the Lewis Dot sy	<b>⊕ \$</b>	
b.	Is the bond NN	[(polar covalent) or (non polar covalent)] (ci	ircle one) (4 pts)
2. Give t	the Lewis Dot Structure	of the following by completing the following	g. H <sub>3</sub> C <sub>2</sub> F O
a. Show	your work for the count $(1\overline{e}) + 2($	for the valence electrons in the entire molecular $(4\overline{e}) + 7\overline{e} + 6\overline{e} = 2$	ules. (6 pts) 24
	4 c	P O	,
	en the following two stru her structure is incorre	actures, choose the correct structure and then et. (6 pts)	explain one reason why
H	CH)	H-H-	O: (3) F cost expand octet-2nd
12×	(a) 2 = 24E	$0/4 \times 2 = 28$	E Comorgo 2200
Extra Credit	(3 pts, ½ pt each blank)	For the following given Lewis Dot Structu	
0 C1 6			
CI Set		·	
a. Numb	per of electrons pairs on	the atom with the * for VSEPRT	
b. Numb	per of lone pairs on the a	tom with the * <u>2e</u> r0'	A 4
c. Geom	netry of the electron pair	s at the atom with the * <u>0 Hahl</u>	dral
d. Geom	netry of the molecule at t	he atom with the * <u>CLahed</u>	lial
e. Bond	angle at the atom with the	he * 90 G	
	=	3 12	

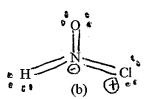
Quiz VII General Chemistry I Lecture Spring 15 Dr. Hahn 20 pts 4/21 T 8:30am form A quiz # 5

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Name	Name (sign_name)	
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Please show all work for full credit and for partial credit. <u>Final Exam 4/30</u> 8:30 am Tues, Thurs class 8:30-10:30am; 9:55 am Tues, Thurs class 3-4pm in LSF 301 1/3 multiple choice, 1/3 short answer, 1/3 long answer

- 1. a. Give the Lewis Dot symbol of the element N (4 pts)
  - b. Number of valence electrons for the element above is: \_\_\_\_\_ electrons (4 pts)
- 2. Give the Lewis Dot Structure of the following by completing the following. HNOCI+1
- a. Show your work for the count of the valence electrons in the entire molecules. (6 pts)
- b. Given the following two structures, choose the correct structure and then explain one reason why the <u>other</u> structure is <u>incorrect.</u> (6 pts)





- a. Number of electrons pairs on the atom with the \* for VSEPRT \_\_\_\_\_
- b. Number of lone pairs on the atom with the \*
- c. Geometry of the electron pairs at the atom with the \*
- d. Geometry of the molecule at the atom with the \*
- e. Bond angle at the atom with the \*
- f. Hybridization at the atom with the \*

Name	Name
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Please show all work for full credit and for partial credit. <u>Final Exam 4/30</u> 8:30 am Tues, Thurs class 8:30-10:30am; 9:55 am Tues, Thurs class 3-4pm in **LSF 301** 1/3 multiple choice, 1/3 short answer, 1/3 long answer

- 1. a. Give the Lewis Dot symbol of the element Cl (4 pts)
  - b. Number of valence electrons for the element above is: \_\_\_\_\_ electrons (4 pts)
- 2. Give the Lewis Dot Structure of the following by completing the following. H<sub>3</sub>N C Cl<sub>2</sub>
- a. Show your work for the count of the valence electrons in the entire molecules. (6 pts)
- b. Given the following two structures, choose the correct structure and then explain one reason why the <u>other</u> structure is <u>incorrect</u>. (6 pts)



- a. Number of electrons pairs on the atom with the \* for VSEPRT \_\_\_\_\_
- b. Number of lone pairs on the atom with the \*
- c. Geometry of the electron pairs at the atom with the \*
- d. Geometry of the molecule at the atom with the \*
- e. Bond angle at the atom with the \*
- f. Hybridization at the atom with the \*

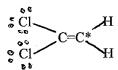
	<b>3</b> T
Name	Name
(print name)	(sign name)

Please show all work for full credit and for partial credit. <u>Final Exam 4/30</u> 8:30 am Tues, Thurs class 8:30-10:30am; 9:55 am Tues, Thurs class 3-4pm in **LSF 301** 1/3 multiple choice, 1/3 short answer, 1/3 long answer

(4 pts)

- 1. a. Give the Lewis Dot symbol of the element C
  - b. Is the bond N----O [(polar covalent) or (non polar covalent)] (circle one) (4 pts)
- 2. Give the Lewis Dot Structure of the following by completing the following. H<sub>2</sub>C S O
- a. Show your work for the count of the valence electrons in the entire molecules. (6 pts)
- b. Given the following two structures, choose the correct structure and then explain one reason why the <u>other</u> structure is <u>incorrect</u>. (6 pts)





- a. Number of electrons pairs on the atom with the \* for VSEPRT \_\_\_\_\_
- b. Number of lone pairs on the atom with the \*
- c. Geometry of the electron pairs at the atom with the \*
- d. Geometry of the molecule at the atom with the \*
- e. Bond angle at the atom with the \*
- f. Hybridization at the atom with the \*

Quiz VII General Chemistry I Lecture Spring 15 Dr. Hahn 20 pts 4/21 T 9:55 am form B quiz # 2

Name Name (print name) (sign name)

Please show all work for full credit and for partial credit. <u>Final Exam 4/30</u> 8:30 am Tues, Thurs class 8:30-10:30am; 9:55 am Tues, Thurs class 3-4pm in **LSF 301** 1/3 multiple choice, 1/3 short answer, 1/3 long answer

- 1. a. Give the Lewis Dot symbol of the element Se (4 pts)
  - b. Is the bond N----N [(polar covalent) or (non polar covalent)] (circle one) (4 pts)
- 2. Give the Lewis Dot Structure of the following by completing the following. H<sub>3</sub>C<sub>2</sub>F O
- a. Show your work for the count of the valence electrons in the entire molecules. (6 pts)

b. Given the following two structures, choose the correct structure and then explain one reason why the <u>other</u> structure is <u>incorrect</u>. (6 pts)

- a. Number of electrons pairs on the atom with the \* for VSEPRT \_\_\_\_\_
- b. Number of lone pairs on the atom with the \*
- c. Geometry of the electron pairs at the atom with the \*
- d. Geometry of the molecule at the atom with the \*
- e. Bond angle at the atom with the \*
- f. Hybridization at the atom with the \*