Oam Form A quiz#
me of your handwriting.)
etc.) (2 pts)
e format. (2 pts)
n the format shown. The structure which you need to
if drew
s + 5p in
a Volker /

Quiz VII General Chemistry I Lecture	Spring 14 Dr. Hahn 20 pts 4/15 T 8:30am Form A quiz #
Name CG	Name
(print name)	Name (sign name) (I can't read some of your handwriting.)
1. a. Give the electron configuration	of Sn in the format $(1s^2, 2s^2, \dots \text{ etc.})$ (2 pts)
152, 252, 2p6, 352, 4p6, or [kr] 552, 4d"	452, 3d1, 4p6, 552, 4d1, 5p2
b. Give the <u>valence</u> electron confi	guration of the same element in the same format. (2 pts)
55 ² , 5p ²	
	guration diagram for the same element in the format shown. The
· , -	and has no information about the correct structure which you need to
draw. (2 pts) $\frac{1 \nu}{1 \text{ s}} \frac{1 \nu}{2 \text{ s}} \frac{1}{2 \text{ p}}$	1 1 St OK if drew 55 + 5p in same level-
	same level-
2. Periodic Relationships: Circle the	e-one correct choice. (4 pts)
[(As) or (N)] has the larger atom size	e (circle one)
(O) or (Se)] has the higher electroneg	
3. Give the Lewis Dot Symbol for t	
Pin group 5-	has 5 dots ve the number of total valence electrons. Show work. (3 pts)
	-3(1)=14= H grow #1 1 Se grow #6
H se C	
b. Given the two structures drawn Dot structure (3 pts)	n for the formula given above, circle the letter under the correct Lewis
••	the can only
H——Se	H have
	H can only H have H duet
H H	\mathcal{H}
(a) 11 1x2= 14E	(b) ~ ~ ~ ~ ~ ~ ~ ~ ~ ~ ~ ~ ~ ~ ~ ~ ~ ~ ~
Extra Credit: Give the dipole moment	19X2=18E-600 range tarrow (+>) for the bond shown below. (3 pts) more EN than H (Husone) Enasc)
H0	mone En than H (Husone)
+	= (Enasc)
+> =	= (Enasc)

Quiz VII General Chemi	stry I Lecture Spring 14 Dr. Hahn 20 pts 4/15 T 8:30am Form B quiz #
Name K &	Name
Name (print name)	(sign name) (I can't read some of your handwriting.)
1. a. Give the electron of 15^2 , 25^2 ,	configuration of N in the format $(1s^2, 2s^2, \text{ etc.})$ (2 pts)
,	
••••	electron configuration of the same element in the same format. (2 pts)
drawing below is only to draw. (2 pts)	electron configuration diagram for the same element in the format shown. The show format and has no information about the correct structure which you need to 1 1 2 1
1s 2s 2p	25
O Deste de Detailemente	
	os: Circle the one correct choice. (4 pts)
<u> </u>	naller atom size (Civile dne)
[(Al) or (C())] has the lo	ower electronegativity (circle one)
3. Give the Lewis Dot S	Symbol for the element Br (4 pts)
br in gre	mp7 - has ndots
4. a. For the formula 2 (e) +	H ₃ COH give the number of total valence electrons. Show work. (3 pts) $4\overline{e} + 6\overline{e} + 1\overline{e} = /4\overline{e}$ $C \qquad O \qquad H$
h Cirron tha tree at	nictions design for the formille outles about a siral the letter under the correct Letter.
Helemen	pts) H has more translated Lower has more H
H C 0	H: Char octet - rot H O H
(6.1)	= 18e marge (b) 1x2= 14e correst # e
Extra Credit: Give the d	ipole moment arrow (+>) for the bond shown below. (3 pts)
→ (\\ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \ \	Vis less ENtra O)

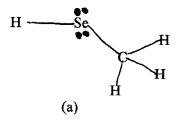
		gold
Quiz VII Ger	neral Chemistry I Lecture Spring 14	Dr. Hahn 20 pts 4/15 T 9:55am Form Å quiz #
Name	Kely	Name (sign_name) (I can't read some of your handwriting.)
(print name)		(sign name) (I can't read some of your handwriting.)
152	And 452, 3010 4pt	in the format $(1s^2, 2s^2, \text{ etc.})$ (2 pts) 3 \mathcal{A}^{0} , 4 \mathcal{P}^{4} or 3 the same element.in the same format. (2 pts)
	452, 484	
	w is only to show format and has no	agram for the same element in the format shown. The information about the correct structure which you need to $\frac{1}{4p}$
normalm	the electron configuration for the 101 $(3-3)$ $(5^2, 25^2, 26^6, 3)$	(in the format 1s, 2s, 2p) (4 pts)
	the Lewis Dot Symbol for the elem	
	The formula $As Cl_2 H$ give the \underline{m} $5e + (95)^2 + 16 = 2$ $As Cl H$	mber of total valence electrons. Show work. (3 pts)
	n the two structures drawn for the fo	rmula given above, circle the letter under the correct Lewis
CI — As=	H. Have chet	CI As H
. •	2=18ē fewé	(b) 10×2= 2Ē
Extra Credit:		>) for the bond shown below. (3 pts)
C N	Cisless	EN GONN)

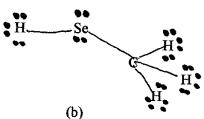
Quiz VII General Chemistry I Lec	cture Spring 14 Dr. Hahn 20 pts 4/15 T 9:55am Form B quiz #
Name Kly	Name (sign_name) (I can't read some of your handwriting.)
(print name)	
)	tion of CI in the format (1s ² , 2s ² , etc.) (2 pts)
15t, 25, 46,3	s2, 30 5 or [Ne] 352, 3p5
b. Give the valence electron c	configuration of the same element in the same format. (2 pts)
35h 385	
	configuration diagram for the same element in the format shown. The mat and has no information about the correct structure which you need to
1s 2s 2p	1V 3P
2 Give the electron configura	ation for the ion Se^{-2} (in the format $1s^2$, $2s^2$, $2p^2$) (4 pts)
normal Se e config-	TAD 452, 3d10, 4p4
Se ⁻² $(+2\bar{e})$ $$ 3 Give the Lewis Dot Symbol	$(Ar) 45^{2}, 3d^{10} 4p^{6}$ of for the element • Ge•(4 pts)
Ge in avour 4	
Ge in group 4 has 4 dots	S S S S S S S S S S S S S S S S S S S
, ,	F give the number of total valence electrons. Show work. (3 pts)
	+ 6ē + 7ē = 18ē
H C	9 F
b. Given the two structures d	frawn for the formula given above, circle the letter under the correct Lewis
Dot structure (3 pts)	(Period 27 lower element
.• •	has more than ottet
H - C-F	Has more than other O has loe + C has lo notallowed
(a) 9 x 2 = 18 E	(b)
,	11x2=22ē
Extra Credit: Give the dipole mor	ment arrow (+>) for the bond shown below. (3 pts)
C Br	c is less EN Gran Br
+	

- 1. a. Give the electron configuration of
- Sn in the format $(1s^2, 2s^2, \dots \text{ etc.})$ (2 pts)
- b. Give the <u>valence</u> electron configuration of the same element in the same format. (2 pts)
- Give the valence electron configuration diagram for the same element in the format shown. The drawing below is only to show format and has no information about the correct structure which you need to draw. (2 pts)

$$\frac{1}{1s} \quad \frac{1}{2s} \quad \frac{1}{2p} \quad -$$

- 2. Periodic Relationships: Circle the one correct choice. (4 pts)
- [(As) or (N)] has the larger atom size (Civcle One)
- [(O) or (Se)] has the higher electronegativity (CIVULONE)
- 3. Give the Lewis Dot Symbol for the element (4 pts)
- 4. a. For the formula HSeCH₃ give the number of total valence electrons. Show work. (3 pts)
 - b. Given the two structures drawn for the formula given above, circle the letter under the correct Lewis Dot structure (3 pts)





Extra Credit: Give the dipole moment arrow (+---->) for the bond shown below. (3 pts)

H ----- O

- b. Give the <u>valence</u> electron configuration of the same element in the same format. (2 pts)
- c. Give the <u>valence</u> electron configuration diagram for the same element in the format shown. The drawing below is only to show format and has no information about the correct structure which you need to draw. (2 pts)

$$\frac{1 V}{1s} \quad \frac{1 V}{2s} \quad \frac{1}{2p} \quad \dots$$

2. Periodic Relationships: Circle the one correct choice. (4 pts)

- 3. Give the Lewis Dot Symbol for the element Br (4 pts)
- 4. a. For the formula H₃COH give the <u>number of total valence electrons</u>. Show work. (3 pts)
 - b. Given the two structures drawn for the formula given above, circle the letter under the correct Lewis Dot structure (3 pts)

Extra Credit: Give the dipole moment arrow (+---->) for the bond shown below. (3 pts)

N ----- O

- 1. a. Give the electron configuration of
- Se in the format $(1s^2, 2s^2, \dots \text{ etc.})$ (2 pts)
- b. Give the valence electron configuration of the same element in the same format. (2 pts)
- Give the valence electron configuration diagram for the same element in the format shown. The drawing below is only to show format and has no information about the correct structure which you need to draw. (2 pts)

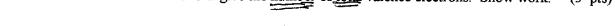
$$\frac{1}{1s} \frac{1}{2s} \frac{1}{2p}$$

- Mg^{+2} (in the format 1s², 2s², 2p²...) (4 pts) 2 Give the electron configuration for the ion
- Give the Lewis Dot Symbol for the element 3 Mg (4 pts)
- a. For the formula As Cl₂ H give the <u>number of total</u> valence electrons. Show work. (3 pts)
 - b. Given the two structures drawn for the formula given above, circle the letter under the correct Lewis Dot structure (3 pts)



Extra Credit: Give the dipole moment arrow (+---->) for the bond shown below. (3 pts)

C ---- N



b. Given the two structures drawn for the formula given above, circle the letter under the correct Lewis Dot structure (3 pts)



Extra Credit: Give the dipole moment arrow (+---->) for the bond shown below. (3 pts)

C ------ Br