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Quiz III General Chemistry I Lecture Spring 14 Dr. Hahn 20 pts 2/11 T 8:30am Form A quiz #
Name Name (print name) (I can't read some of your handwriting.)
Please show all work for full credit. If you show work you may also get partial credit.
1. From the list of molecules shown below circle all <u>ionic</u> compounds (5 pts)  SO <sub>2</sub> KCl SF <sub>6</sub> PBr <sub>3</sub> CaSO <sub>4</sub> ( fax in periodic)  2. Write the correct formula for the following ionic compound. If you show work you may earn some partial credit.
2. Write the correct formula for the following ionic compound. If you show work you may earn some partial credit.
Mg and P charge on Mg $+2$ charge on P $-3$ (1 pt each, 2 pts total)  Correct formula is  Correct formula is
correct formula is(2 pts)
3a. Is the following formula an [ (empirical formula) or (molecular formula)] (circle one) (3 pts)?
C4H10 +2 = C2 H5 has comma denominates
b. Given the following structure, give the name of the polyatomic ion:
CH3COO acetate (also C2 H3O2) (2 pts)
c. Given the name of the polyatomic ion, give the structure (2 pts)
phosphate $P04^{-3}$
4. Name the following molecule: Mg(NO3)2 (4 pts)  bihary i Onic - ho # prefix  mgg he sillen hitrate
Extra Credit: What is its molecular weight (formula mass) for the compound. Show work. (2 pts)
$Na_2SO_4$ 2(23,0) + 32,07 + 4(16,0) = 142,019
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Quiz III General (	Chemistry I Lecture Spring 14	Dr. Hahn 20 pts 2/11 T 8:30am Form	nBquiz#
Name (	Dy	37	
(print name)		_ Name (sign_name) (I can't read some of your	handwriting )
Please show all wo	ork for full credit. If you show	work you may also get partial credit.	audewitting.)
		circle all covalent compounds (5 pts)	
SO <sub>2</sub> KCl  2. Write the conceptral cred	SF <sub>6</sub> PBr <sub>3</sub> Ca SO <sub>2</sub> orrect formula for the following	g ionic compound. If you show work you  + (-1) M3 = 0	on metal u may earn some
Ca and NO <sub>3</sub>	charge on Ca +2	charge on NO <sub>3</sub> (1 pt each,	2 pts)
correct formula is_	Ca (NO3)2	(2 pts)	
3a Is the follow $SF_6$	ving formula an [empirical fo	ormula) or (molecular formula)] (circle or	ne) ? (3 pts)
b. Given the fo	in houte	ne of the polyatomic ion: (2 pts)	
c. Given the na	ame of the polyatomic ion, give	the structure (2 pts)	
Sulfate	5.04		
4 Name the fold the Air Sul	llowing molecule: S2F4  me + A A An  fur Eetra	(4 pts) Covalent - # ne - ending + ide) fluoride	e protis
Extra Credit: What Al(NO <sub>3</sub> ) <sub>3</sub>	is its molecular weight (formuly $0.7, 0 + (4, 0)$ )	la mass) for the compound. Show work. $ \begin{array}{cccccccccccccccccccccccccccccccccc$	(2 pts) 3= 213,03g mal

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Quiz III General Chemistry I Lecture Spring 14 Dr. Hahn 20 pts 2/11 T 9:55 am Form A quiz #
NameNameName
(print name) (I can't read some of your handwriting.)
Please show all work for full credit. If you show work you may also get partial credit.
1. From the list of molecules shown below circle all ionic compounds (5 pts)
1. From the list of molecules shown below circle all <u>ionic</u> compounds (5 pts)  NO <sub>2</sub> (Mg(NO <sub>3</sub> ) <sub>2</sub> ) NCl <sub>3</sub> (BaO) SeBr <sub>2</sub> for in Periodic table  2. Write the correct formula for the following ionic compound made up of the two parts given below.  If you show work you may earn some partial are differential.
if you show work you may earn some partial cledit.
Na and $SO_4^{-2}$ charge on Na charge on $SO_4$ $O(4)$ charge on $SO_4$ $O(4)$ $O(4)$
Na 2 504)
correct formula is (2 pts)
3a. the number prefix for the number 5 is 10 A (3 pts)
b. Given the following structure, give the name of the polyatomic ion (2 pts)
NH4 ammonium
c. Given the name of the polyatomic ion, give the structure (2 pts)
Acetate CH3 COO (also OK)  4. Name the following molecule: Na2O (4 pts) i Mic - Moth  Sodium & Kide Pretix  Element name - ending + ide  & X14 Sen 1+ ide
Extra Credit: What is its molecular weight (formula-mass) for the compound. Show work. (2 pts)
(NH <sub>4</sub> ) <sub>2</sub> 0 ((14,01) # 2) + ((1,01) # 4#2) + (6.0 = 52,1)

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Quiz III	General Ch	nemistry I Lec	ture Spring 14 Dr.	Hahn 20 pts 2/11	T 9:55 am Form B quiz#	-
Name _		ey	N	ame		
(print na	me)	0	(si	gn name) (I can't i	read some of your handwriting.)	
Please sl	now all wor	k for full credi	t. If you show wor	k you may also get	partial credit.	
1. ]	From the lis	t of chemical t	formulas shown be	ow circle all coval	ent compounds (5 pts)	1.0
$(NO_2)$	Mg(NO <sub>3</sub> )		BaO SeBr <sub>2</sub>	Close Lable	Egether in per-	100010 
2. V sl	Vrite the cor how work y	rect formula f ou may earn s	or the ionic compo ome partial credit.	und made up of the	e two parts given below. If you	
Al an	d 0 +3)	charge Charge	ge on Al	charge on C	1 pt each, 2 pts)	
correct fo	ormula for th	ne ionic comp	ound is M2	03	_ (2 pts)	
3a. th	e number p	refix for the n	umber 8 is	octa	(3 pts)	
b. G	iven the foll	lowing structu	re, give the name o	f the polyatomic io	n:	>
PO <sub>4</sub> -3	P	osph	ate	(2	pts) # pretixt Same as	
c. G	iven the nan	ne of the polya	atomic ion, give the	structure (2 pts)	Same	
Nitrate _		<u>0 z</u>		,	biras	
4. Na	ame the foll	owing molecu OSPLOY	le: $P_2S_5$ (4 pts) $M_2S_3 = M_2S_3$	-coval	SWI fux I + id	2
2 <sup>C</sup>	li Rh	0 gho)	rus per	tasul	ficle	
Extra-Cree	dit. What i	A Mar	weight (formula n	element	none-enligtic	de)
Mg <sub>3</sub> (PO <sub>4)</sub>	2	24.31	) * 3) +	[60,92	)* 2] + (16.0)*(4)*	4(2)
		mg		P		
	2	262	873/mi	l		

	Name
(print name)	Name (sign_name) (I can't read some of your handwriting.)
Please show all work for full cred	lit. If you show work you may also get partial credit.
1. From the list of molecules	s shown below circle all <u>ionic</u> compounds (5 pts)
SO <sub>2</sub> KCl SF <sub>6</sub> PE	Br <sub>3</sub> CaSO <sub>4</sub> .
Write the correct formula f partial credit.	for the following ionic compound. If you show work you may earn some
Mg and P charge on M	Mg charge on P (1 pt each, 2 pts total)
correct formula is	(2 pts)
3a. Is the following formula a	an [ (empirical formula) or (molecular formula)] (circle one) (3 pts)?
$C_4H_{10}$	
-	ure, give the name of the polyatomic ion:
CH3COO"	(2 pts)
	vatamia ion, give the atmeture (2 nta)
c. Given the name of the poly	rationne fon, give the structure (2 pts)
c. Given the name of the polyochosphate	
phosphate	
phosphate	
phosphate	
ohosphate4. Name the following molecu	

Quiz 1	III Geņeral Chemistry	I Lecture S	Spring 14 Dr. Hab	m 20 pts 2/1	1 T 8:30am	Form B quiz	#_{
Name			Name				
	name)		(sign n	ame) (I can't	read some of	your handwritin	ng.)
Please	show all work for fu	l credit. If y	ou show work yo	u may also ge	et partial credi	t.	
1.	From the list of mo	ecules show	n below circle all	covalent con	npounds (5 pt:	s)	
$SO_2$	KCl SF <sub>6</sub>	PBr <sub>3</sub>	Ca SO <sub>4</sub>		•		
2.	Write the correct for partial credit.	mula for the	following ionic c	ompound. If	f you show wo	ork you may ear	n some
Са	and NO <sub>3</sub> charg	e on Ca	charg	e on NO <sub>3</sub>	(1 pt	each, 2 pts)	
correc	t formula is		(	(2 pts)			
3a b.	Is the following form $SF_6$ Given the following		,	·	V. V.	rele one) ? (3 pt	s)
NO <sub>3</sub>			•				
c.	Given the name of th	e polyatomi	c ion, give the stru	ecture (2 pts)			
Sulfate	<u> </u>				·		
	Name the following 1	nolecule:	S <sub>2</sub> F <sub>4</sub> (4 pts)				
Extra C	Credit: What is its mo	olecular weig	ght (formula mass`	) for th <del>e</del> com	oound. Show	work. (2 pts)	
AI(NO:		•	-	1	•	<b>、</b> 1 <i>)</i>	

Quiz III General Chemistry I Lecture Spring 14 Dr. Hahn 20 pts 2/11 T 9:55 am Form A quiz #
Name Name (print name) (I can't read some of your handwriting.)
(print name) (sign name) (I can't read some of your handwriting.)
Please show all work for full credit. If you show work you may also get partial credit.
1. From the list of molecules shown below circle all <u>ionic</u> compounds (5 pts)
$NO_2$ $Mg(NO_3)_2$ $NCl_3$ $BaO$ $SeBr_2$
<ol><li>Write the correct formula for the following ionic compound made up of the two parts given below. If you show work you may earn some partial credit.</li></ol>
Na and $SO_4^{-2}$ charge on Na charge on $SO_4$ (1 pt each, 2 pts)
correct formula is(2 pts)
3a. the number prefix for the number 5 is (3 pts)
b. Given the following structure, give the name of the polyatomic ion (2 pts)
NH <sub>4</sub> <sup>+</sup>
c. Given the name of the polyatomic ion, give the structure (2 pts)
Acetate
4. Name the following molecule: Na <sub>2</sub> O (4 pts)
Extra Credit: What is its molecular weight (formula mass) for the compound. Show work. (2 pts)
$(NH_4)_2O$

Quiz III General Chemistry I Lecture Spring	14 Dr. Hahn 20 pts 2/11 T 9:55 am Form B quiz#
Name	Name
(print name)	Name (sign name) (I can't read some of your handwriting.)
Please show all work for full credit. If you sho	ow work you may also get partial credit.
1. From the list of chemical formulas sho	own below circle all covalent compounds (5 pts)
NO <sub>2</sub> Mg(NO <sub>3</sub> ) <sub>2</sub> NCl <sub>3</sub> BaO	SeBr <sub>2</sub> .
2. Write the correct formula for the ionic show work you may earn some partial of	compound made up of the two parts given below. If you credit.
Al and O charge on Al	charge on O(1 pt each, 2 pts)
correct formula for the ionic compound is	(2 pts)
3a. the number prefix for the number 8	is(3 pts)
b. Given the following structure, give the r	name of the polyatomic ion:
PO <sub>4</sub> -3	(2 pts)
c. Given the name of the polyatomic ion, g	give the structure (2 pts)
Nitrate	
4. Name the following molecule: P <sub>2</sub> S <sub>5</sub> (	4 pts)
Extra Credit: What is its molecular weight (for	mula mass) for the compound. Show work. (2 pts)

 $Mg_3(PO_4)_2$