_	nic Chemistry I (CHEM 340) Fall 15 Dr. Hahn MWF 11 am Quiz I Form A Exam#
	Name Print Name
Please	e show work on all questions for partial credit even on questions which do not specify. (25 total pts)
1.	a. For the element Se show the electron configuration for all electrons in the format $1s^2$ $2s^2$
etc (2	2pts each letter, 12 pts total) $(52, 25^2, 2p^2, 35^2, 3p^6, (-7pA)$
4	2pts each letter, 12 pts total) (52, 252, 2pb, 352, 3pb, Capt)  For the same element show the electron configuration for all valence electrons in the same format.
b. 1	For the same element show the electron configuration for all valence electrons in the same format.
c.	For the same element, show the valence electron configuration orbital diagram in the format. Way
(Ab	
ls	
d.	For the same element, what is the group number? e. what is the atomic mass?
	$8.9k$ f. what is the atomic number? $\frac{54}{}$
2.	Give the symbol of one element which belongs to the "s block" of the periodic
1 1	Ca_(2 pts) (any group 1 A + 2 A element)
table.	
3 under	Given the following Lewis dot structures, which is the correct Lewis Dot structure? Circle the letter the <u>correct structure</u> and then give at <u>one reason</u> why the <u>wrong structure is incorrect</u> . (9 pts, 5 pts
	e, 4 pts explanation)  1 H-N-CH3  H-N-CH3  H-N-CH3  CH3  CH3  CH3  CH3  CH3  CH3  CH3
	Then duet, (more
	11 - Then ortet
	10 H H D 20 1 2 (10) + 5 (10)
	a (1) there = 9ē+2(4ē)+5(1ē),
	(EN) (B) (B) N C H
đ	Circum the Salverine Leville Det structure de VSERRT
4.	Given the following Lewis Dot structure, do VSEPRT Q has $20 \times 2 = 40 = 6$
	See has 12 vs
	(C) * (C)
a.	Number of VSEPRT electron pairs around the atom with the * (1 pt)
b.	Number of lone pairs around the atom with the * $2$ (1 pt)

		1.(CHEM 340) Fall 15	Dr. Hahn	MWF 11 am	Quiz I	Form B Ex	xam #
_	Name	Kery		Print Name			b/w
Pleas	se show work o	on all questions for partia	l credit ever	on questions w	hich do no	ot specify. (25	total pts)
c. { 1v ls d.	For the same	e element show the elect e element, show the vale using up and down arrow	25 <sup>2</sup> , 2 Pron configur properties of the control	ation for all value configuration of the configurat	3p 1frence electr D — I rbital diagn	45 <sup>2</sup> ons in the sai	3d <sup>10</sup> me format.
112	\$ <u>.710</u> f.	what is the atomic num	iber ?	0			
choice	Given the for the correct stree, 4 pts explanate $\times 2 = 25$ g. Given the following the following formula $\times 25$ given formula $\times 25$ given formula $\times 25$ g	C Cel  W G Spt  lowing Lewis Dot struct	B to Oures, which one reason c) + (1 e) (1 c) (1 c) (2 c) (3 c) (4 c) (4 c) (4 c) (4 c) (5 c) (6 c) (7 c) (7 c) (8 c) (9 c) (9 c) (9 c) (1	is the correct Le  why the wrons $+3(1e)$	wis Dot st structure = 28 e	ructure? Cir e is incorrect E hora E hora 440	19 pts. 5 pts  19 ts. 5 pts  19 than  20 chest
a.	Number of V	SEPRT electron pairs ar	ound the ato	m with the *	3_(1	pt) = C	Amts as)
b.		ne pairs around the atom		_		•	4
c. ,	VSEPRT mo	lecular geometry for the	atom with th	ne * trig	oral a	slana	(2 pts)
Extra (		the following, give the s					wrong #C or wrong -2 census

Oro	anic Chemistry I (CHEM 340) Fall 15 Dr. Hahn MW 5 pm Quiz I Exam #
	n Name Print Name
	ase show work on all questions for partial credit even on questions which do not specify. (25 total pts)
1. etc 4 b.	a. For the element As show the electron configuration for all electrons in the format $1s^2, 2s^2,$ (2pts each letter, 12 pts total) $15^2, 25^2, 2p^6, 25^2, 3p^6, 45^2, 3d^{10}$ For the same element show the electron configuration for all valence electrons in the same format.
c.	For the same element, show the valence electron configuration orbital diagram in the format:
{ <u>1</u>    s	For the same element, what is the group number? $A = A = A = A = A = A = A = A = A = A $
7	$\frac{492}{}$ f. what is the atomic number? $\frac{33}{}$
2.	Give the symbol of one element which belongs to the "p block" of the periodic
table	1 Original 3A to 8A almot
choid	Given the following Lewis dot structures, which is the correct Lewis Dot structure? Circle the letter on the correct structure and then give at one reason why the wrong structure is incorrect. (9 pts, 5 pts ce, 4 pts explanation)  Compared to the correct structure and then give at one reason why the wrong structure is incorrect. (9 pts, 5 pts ce, 4 pts explanation)  Compared to the correct Lewis Dot structure is incorrect. (9 pts, 5 pts ce, 4 pts explanation)  Compared to the correct Lewis Dot structure is incorrect. (9 pts, 5 pts ce, 4 pts explanation)  Compared to the correct Lewis Dot structure is incorrect. (9 pts, 5 pts ce, 4 pts explanation)  Compared to the correct Lewis Dot structure is incorrect. (9 pts, 5 pts ce, 4 pts explanation)  Compared to the correct Lewis Dot structure is incorrect. (9 pts, 5 pts ce, 4 pts explanation)  Compared to the wrong structure is incorrect. (9 pts, 5 pts ce, 4 pts explanation)  Compared to the correct Lewis Dot structure is incorrect. (9 pts, 5 pts ce, 4 pts explanation)  Compared to the correct Lewis Dot structure is incorrect. (9 pts, 5 pts ce, 4 pts explanation)  Compared to the correct Lewis Dot structure is incorrect. (9 pts, 5 pts ce, 4 pts explanation)  Compared to the correct Lewis Dot structure is incorrect. (9 pts, 5 pts ce, 4 pts explanation)  Compared to the correct Lewis Dot structure is incorrect. (9 pts, 5 pts ce, 4 pts explanation)  Compared to the correct Lewis Dot structure is incorrect.
a.	Number of VSEPRT electron pairs around the atom with the * $\frac{3}{2}$ (1 pt) (= $\frac{2}{2}$ (1 pt) (= $\frac{2}{$
b.	(1 pt)
c.	VSEPRT molecular geometry for the atom with the * trigonul pandra (2 pts)
Gai Gai	Credit: Given the following skeletal structure give the line bond structure which corresponds. (2 pts)

Dlages	Name Print Name e show work on all questions for partial credit even on questions which do not specify. (25 total pts)
1.	a. For the element <b>Se</b> show the electron configuration for all electrons in the format $1s^2$ , $2s^2$ , and $2s^2$ expressed the electron configuration for all electrons in the format $2s^2$ , $2s^2$ , and $2s^2$ expressed the electron configuration for all electrons in the format $2s^2$ , $2s^2$ , and $2s^2$ expressed the electron configuration for all electrons in the format $2s^2$ , $2s^2$
b.	For the same element show the electron configuration for all <u>valence</u> electrons in the same format.
c.	For the same element, show the valence electron configuration orbital diagram in the format:
{ <u>A</u> <b>v</b> Is	etc} using up and down arrows to represent electrons.
d.	For the same element, what is the group number? e. what is the atomic mass?
	f. what is the atomic number?
2.	Give the <b>symbol of one element</b> which belongs to the "s block" of the periodic
table.	(2 pts)
	e, 4 pts explanation)  1 H - N - CH3  H - N - CH3
	10 /C HI  OCH  OCH
4.	Given the following Lewis Dot structure, do VSEPRT  See Co
a.	Number of VSEPRT electron pairs around the atom with the * (1 pt)
	Number of VSEPRT electron pairs around the atom with the * (1 pt)  Number of lone pairs around the atom with the * (1 pt)
a.	Number of VSEPRT electron pairs around the atom with the * (1 pt)

Sign N	Name Print Name
Please	e show work on all questions for partial credit even on questions which do not specify. (25 total pts
1. etc (2	a. For the element $\mathbf{Sn}$ show the electron configuration for all electrons in the format $1s^2$ , 2 epts each letter, 12 pts total)
b.	For the same element show the electron configuration for all <u>valence</u> electrons in the same formation
c.	For the same element, show the valence electron configuration orbital diagram in the format:
{ <u>1</u> V	t 1 etc} using up and down arrows to represent electrons.
d.	For the same element, what is the group number? e. what is the atomic mass?
	f. what is the atomic number?
2.	Give the symbol of one element which belongs to the "d block" of the periodic
table.	(2 pts)
	# (g) # (J) (H) (h)
3.	Given the following Lewis Dot structure, do VSEPRT
a.	Number of VSEPRT electron pairs around the atom with the * (1 pt)
b.	Number of lone pairs around the atom with the *(1 pt)
c.	VSEPRT molecular geometry for the atom with the *(2 g
Extra	Credit: Given the following, give the skeletal structure which corresponds. (2 pts)

Organic Chemistry I (CHEM 340) Fall 15 Dr. Hahn Sign Name					
Please show work on all questions for partial credit even					
1. a. For the element <b>As</b> show the electron co etc (2pts each letter, 12 pts total)	on figuration for all electrons in the format $1s^2, 2s^2,$				
b. For the same element show the electron configur	ration for all <u>valence</u> electrons in the same format.				
c. For the same element, show the valence electron	configuration orbital diagram in the format:				
$\{ \frac{1}{1} $ etc $\}$ using up and down arrows to represent	ent electrons.				
d. For the same element, what is the group number	? e. what is the atomic mass?				
f. what is the atomic number?					
2. Give the <b>symbol of one element</b> which belongs	to the "p block" of the periodic				
table(2 pts)					
Given the following Lewis dot structures, which under the <u>correct structure</u> and then give at <u>one reaso</u> choice, 4 pts explanation)	is the correct Lewis Dot structure? Circle the letter on why the wrong structure is incorrect. (9 pts, 5 pts				
3. Given the following Lewis Dot structure, do VS	EPRT				
a. Number of VSEPRT electron pairs around the at	tom with the * (1 pt)				
b. Number of lone pairs around the atom with the *	*(1 pt)				
c. VSEPRT molecular geometry for the atom with	the *(2 pts)				
Extra Credit: Given the following skeletal structure giv	·				